GROUP # 5

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DRA. ANDOY
Organic Chemistry (Lab)

M9 Check-In Activity
Group Report

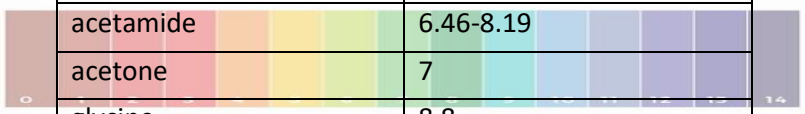
1. Why do we need to calibrate the pH meter before using?

- pH meter need to be calibrated regularly to ensure accuracy.

2. Why is distilled water used in washing the electrodes after testing each sample?

- The purpose of using distilled water is to remove build up residues and impurities.

3. Complete the table.



Test Substance	pH
acetic acid	2.4
monochloroacetic acid	3.05
acetamide	6.46-8.19
acetone	7
glycine	8.8
lysine	10.5
isopropyl alcohol	8.0
phenol	5-6

4. Compare the pH values of the following compounds:

a. Acetic acid and monochloroacetic acid

acetic acid has a pH value of 2.4 while monochloroacetic acid has 3.05. According to their pH value acetic acid is more acidic that monochloroacetic acid due to its lower pH value.

b. Acetamide and acetone

Acetamide varies about 6.46 to 8.19 pH. On the other hand, Acetone has a pH value of 7. Acetamide consider as base because of its pH while given the pH value of acetone it can consider as neutral as acetone is neither acid nor base.

c. Glycine and lysine

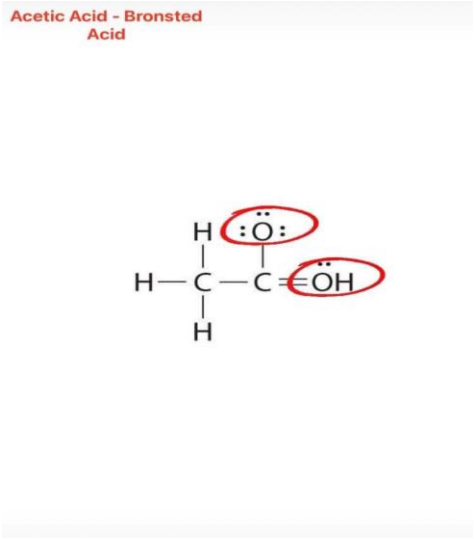
Glycine has a pH value of 8.8 and Lysine has 10.5. Since, Lysine has a higher Ph value it means that it has stronger base while the Glycine has a stronger acid.

d. Isopropyl alcohol and phenol

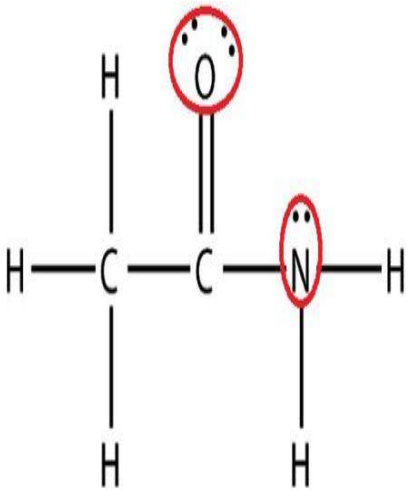
Alcohol has a Ph value of 8.0 which means that it has a stronger base and Phenol has a 5-6 Ph value which means that it has a stronger acid.

5. Which of the following samples above are classified as stated below. Support your answers by writing the structural formula and encircling the part that makes it acidic, basic or neutral as the case may be:

a. Bronsted acid

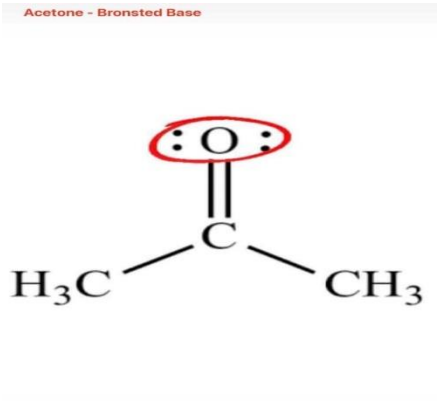


Acetic Acid

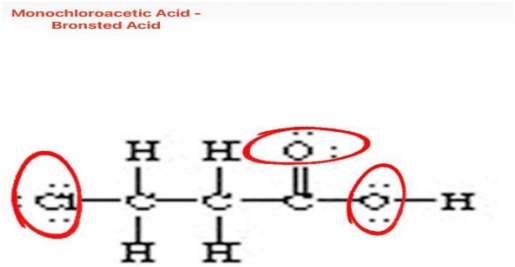


Acetamide

b. Bronsted base

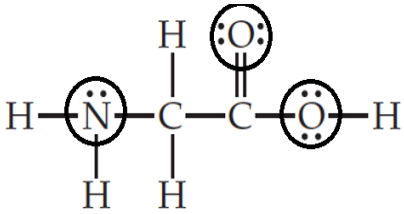


Acetone



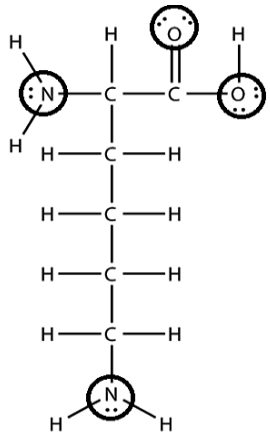
Monochloroacetic acid

c. Lewis acid

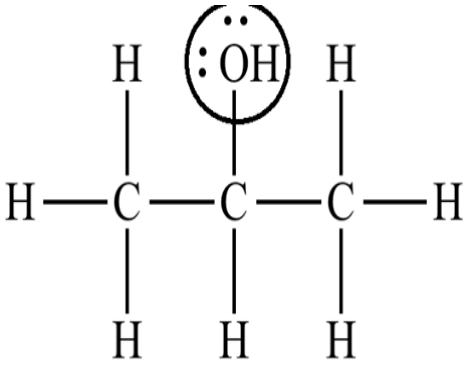


Glycine

d. Lewis base



Lysine



Isopropyl alcohol