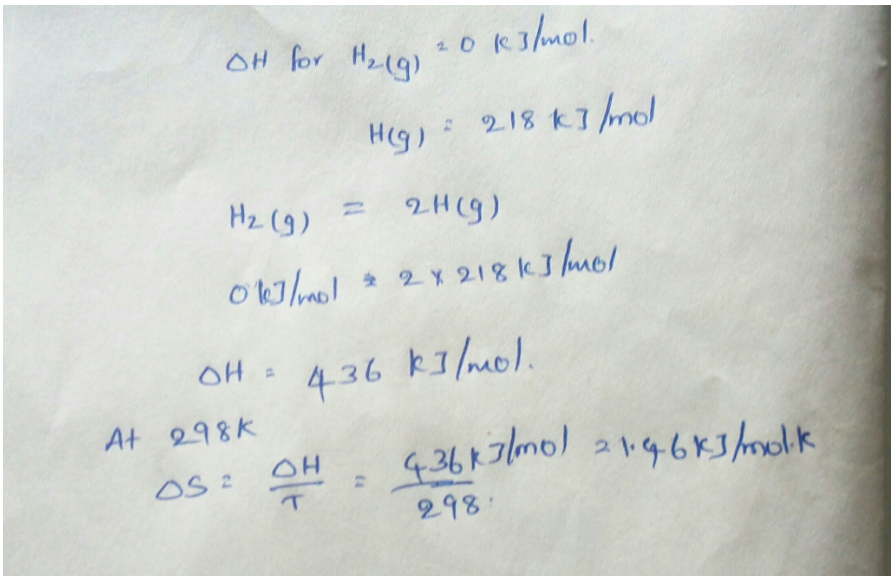
Question: Calculate the standard entropy change for the following reacti...

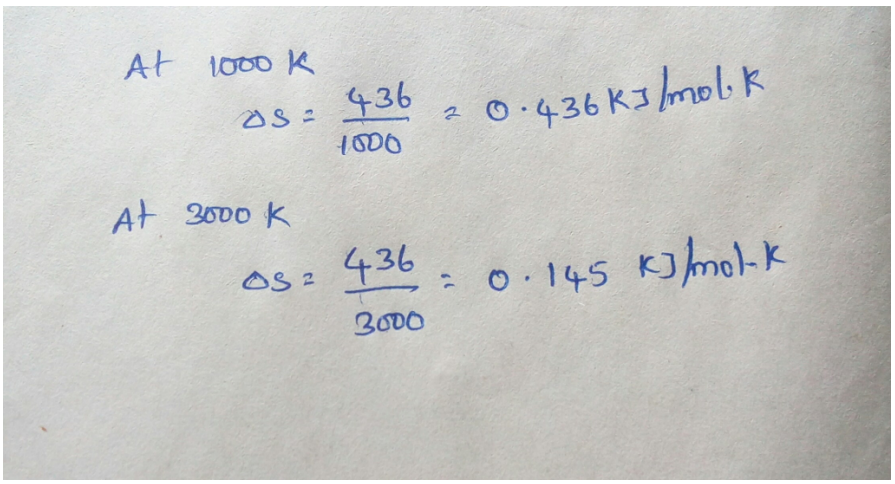
Calculate the standard entropy change for the following reaction H2 (g) = 2H (g) at 298K, 1000 K and 3000k.

Expert Answer

honeymona answered this
1,124 answers

Was this answer helpful? 0 0



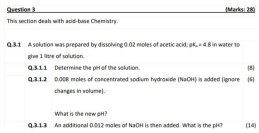


Comment

dissolving 0.02 moles of acetic acid (HOAc: pKa = 4.8)
in water to give 1 liter of solution. 1) What is the pH? 2)

See answer

section deals with acid-base



Question 3 (Marks: 28)
This section deals with acid-base Chemistry.
Q.3.1 A solution was prepared by dissolving 0.02 moles of acetic acid; $pK_a = 4.8$ in water to give 1 litre of solution.
Q.3.1.1 Determine the pH of the solution. (8)
Q.3.1.2 0.008 moles of concentrated sodium hydroxide (NaOH) is added (ignore changes in volume). (6)
What is the new pH?
Q.3.1.3 An additional 0.012 moles of NaOH is then added. What is the pH? (14)

See answer

See more questions for subjects you study

Questions viewed by other students

Q: 2.57 Calculate Δ11° for the dissociation H2(g) = 2H(g) at 0, 298, and 3000 K. The value at 0 K is equal to the spectro-scopic dissociation energy Do

A: See answer

Q: 6.4 Copper exists in the state T-298K·P-1 atm. Calculate the temperature to which the copper must be raised at 1 atm pressure to cause the same increase in molar enthalpy as is caused by increasing its pressure to 1000 atm at 298 K. The molar volume of Cu at 298 K is 7.09 cm3, and the volumetric coefficient of thermal expansion is 0.501x10-4K-1. These values can be taken as being...

A: See answer

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