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Question: Calculate the standard entropy change for the following reacti...

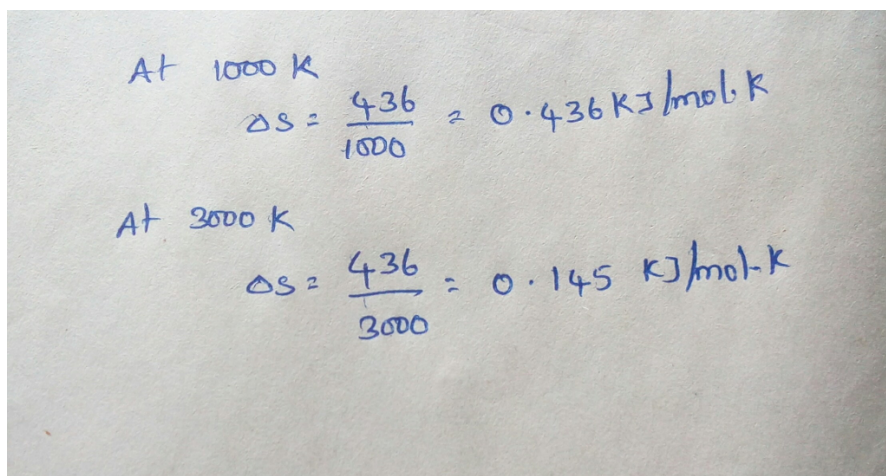
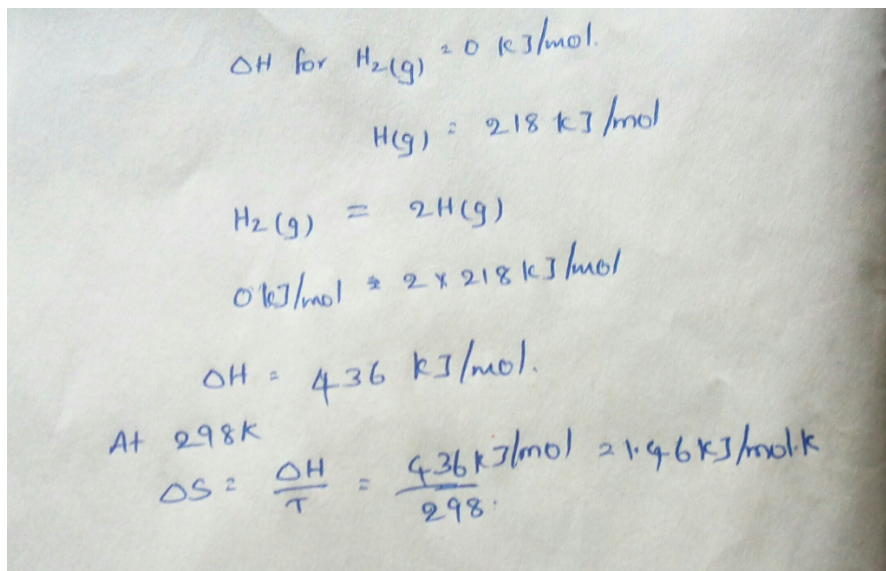
Calculate the standard entropy change for the following reaction $\text{H}_2(\text{g}) = 2\text{H}(\text{g})$ at 298K, 1000 K and 3000K.


Expert Answer



honeymona answered this
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dissolving 0.02 moles of acetic acid (HOAc; $pK_a = 4.8$) in water to give 1 liter of solution. 1) What is the pH? 2)

[See answer](#)

section deals with acid-base

Questions **Marked 2/0**

This section deals with acid-base Chemistry.

Q.3.1 A solution was prepared by dissolving 0.02 moles of acetic acid, $pK_a = 4.8$ in water to give 1 liter of solution.

Q.3.1.1 Determine the pH of the solution. (8)

Q.3.1.2 0.008 moles of concentrated sodium hydroxide (NaOH) is added (ignore changes in volume). (8)

What is the new pH?

Q.3.1.3 An additional 0.02 moles of NaOH is then added. What is the pH? (8)

[See answer](#)

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Q: 2.57 Calculate Δ_{11}° for the dissociation $H_2(g) = 2H(g)$ at 0, 298, and 3000 K. The value at 0 K is equal to the spectroscopic dissociation energy D_0

A: [See answer](#)

Q: 6.4 Copper exists in the state T-298K-P-1 atm. Calculate the temperature to which the copper must be raised at 1 atm pressure to cause the same increase in molar enthalpy as is caused by increasing its pressure to 1000 atm at 298 K. The molar volume of Cu at 298 K is 7.09 cm^3 , and the volumetric coefficient of thermal expansion is $0.501 \times 10^{-4} K^{-1}$. These values can be taken as being...

A: [See answer](#)

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